

RAMAN SPECTROSCOPY OF LASER-MATTER INTERACTIONS

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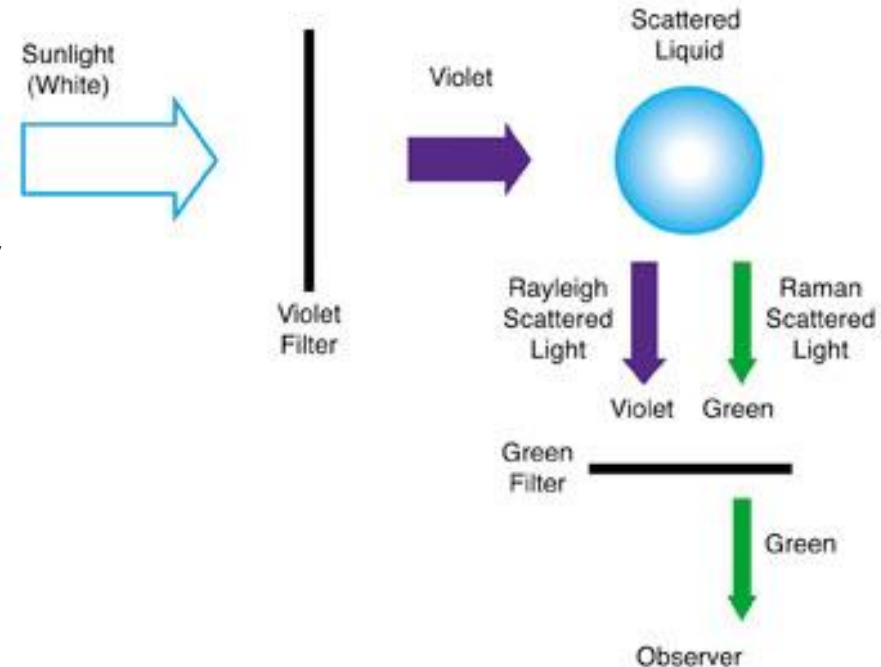
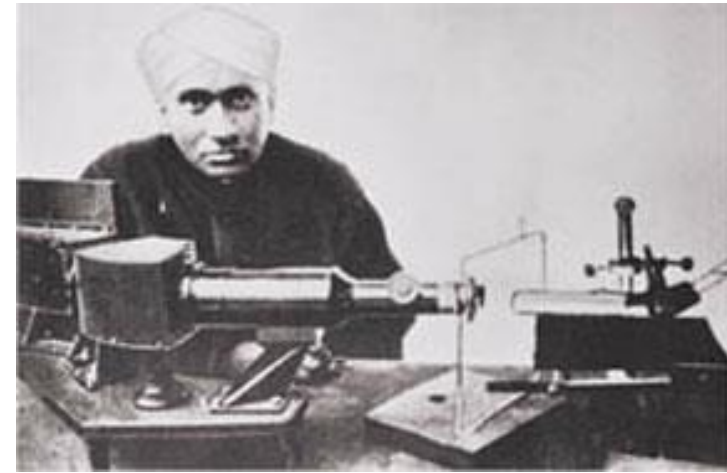
Raman scattering

C. V. Raman in 1928
Nobel Prize, 1930

Investigation of light scattering by water droplets with focused sunlight and filters

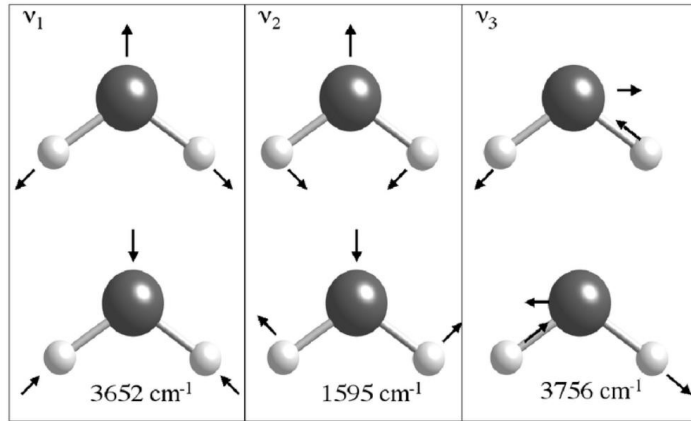
Appearance of a different (green) color in the scattered (violet) light, which had 10^{-7} times lower intensity

G. Landsberg, L. Mandelstam
Combinatorial scattering in crystals

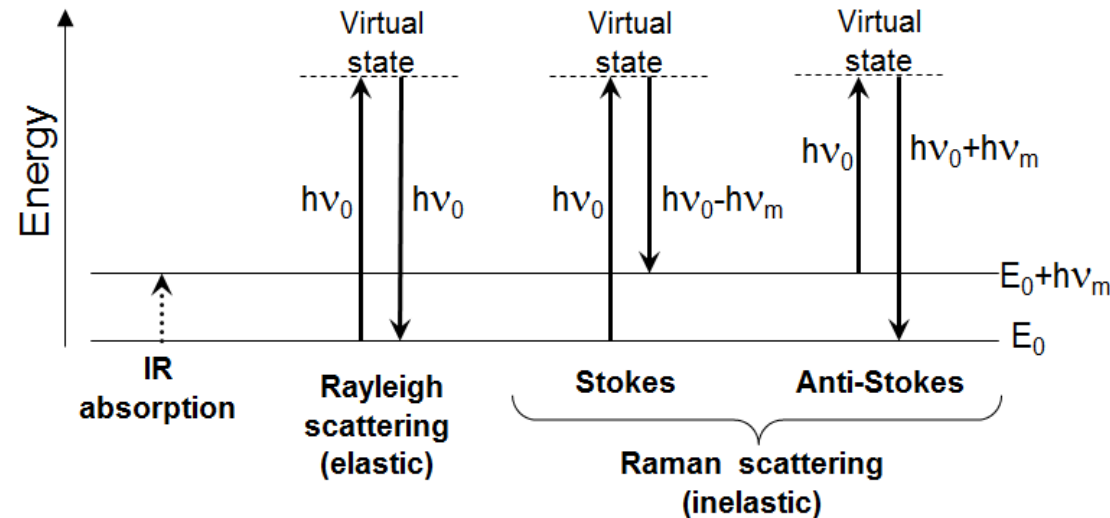


Raman scattering

Inelastic light scattering due to the interaction of the incident light with matter.



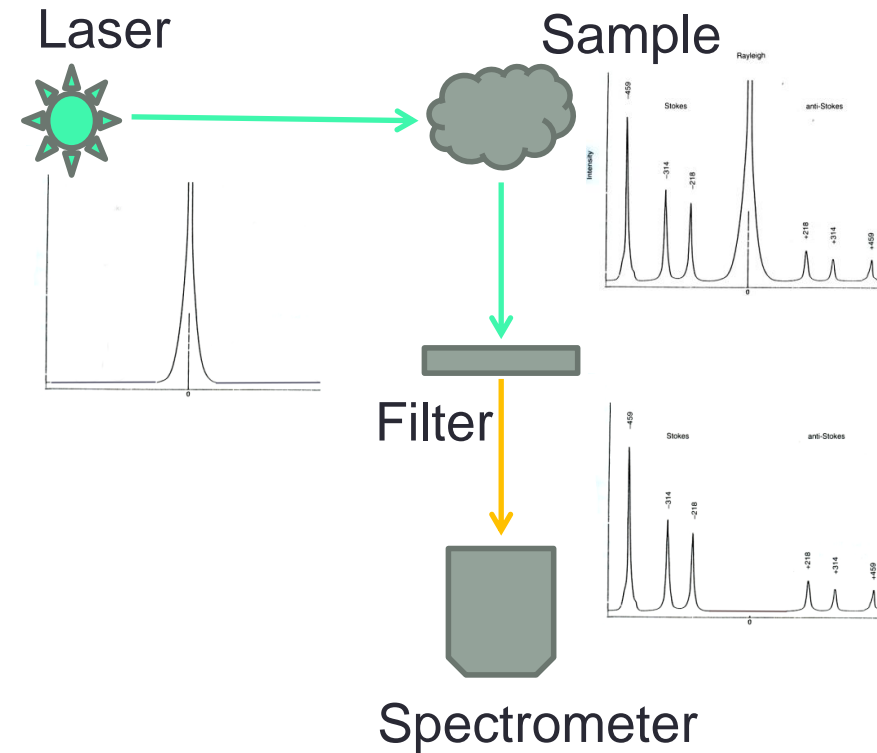
Vibrations of a water molecule



The energy difference of the excitation and Raman scattered light equals to the energy of a normal vibration of the medium (molecule, crystal).

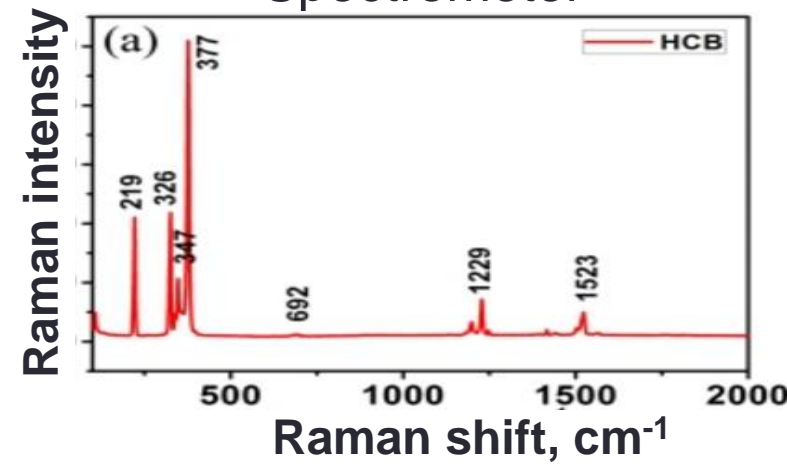
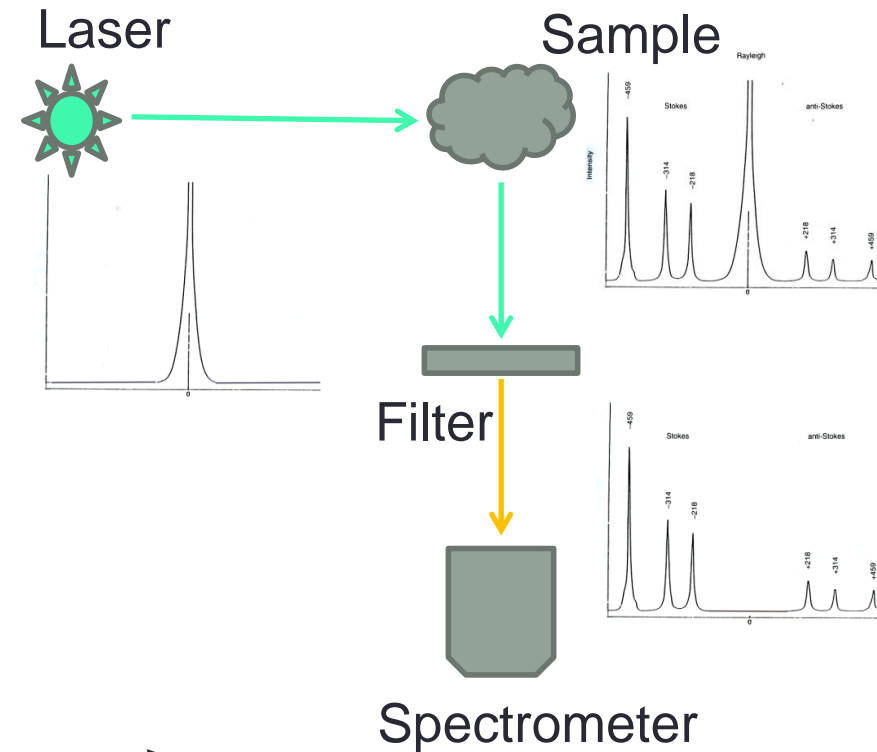
Raman spectroscopy

- Inelastic light scattering
- Excitation with a monochromatic light source (laser)
- Recording the spectrum of the scattered light in wavelength region different from the excitation



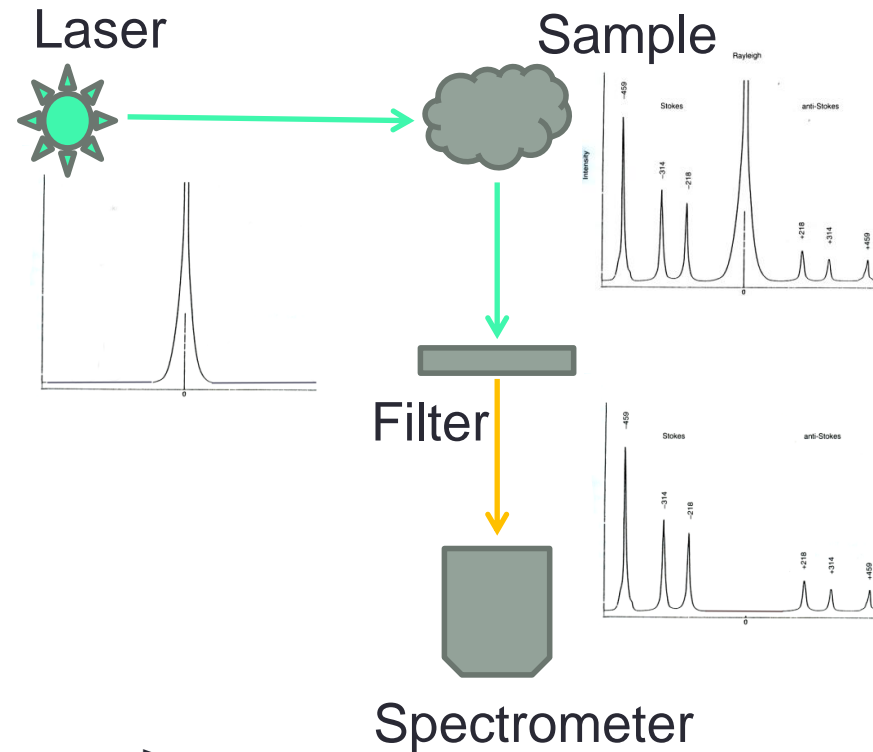
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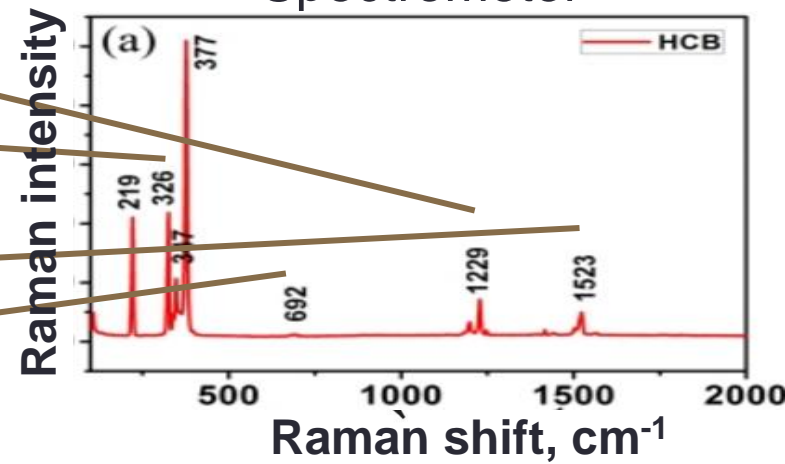
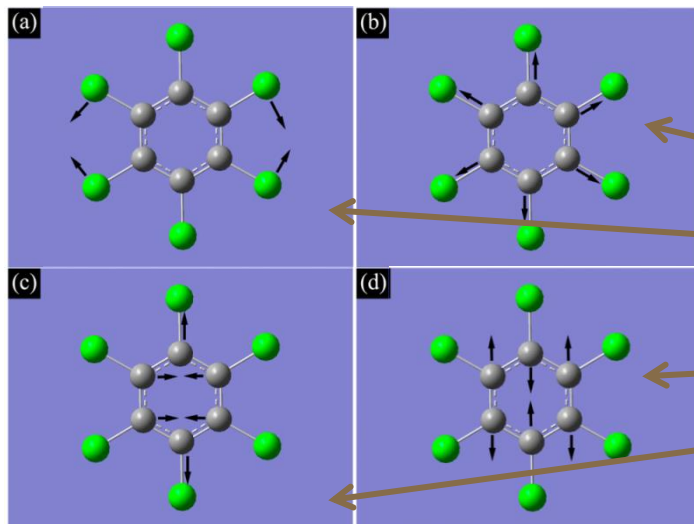


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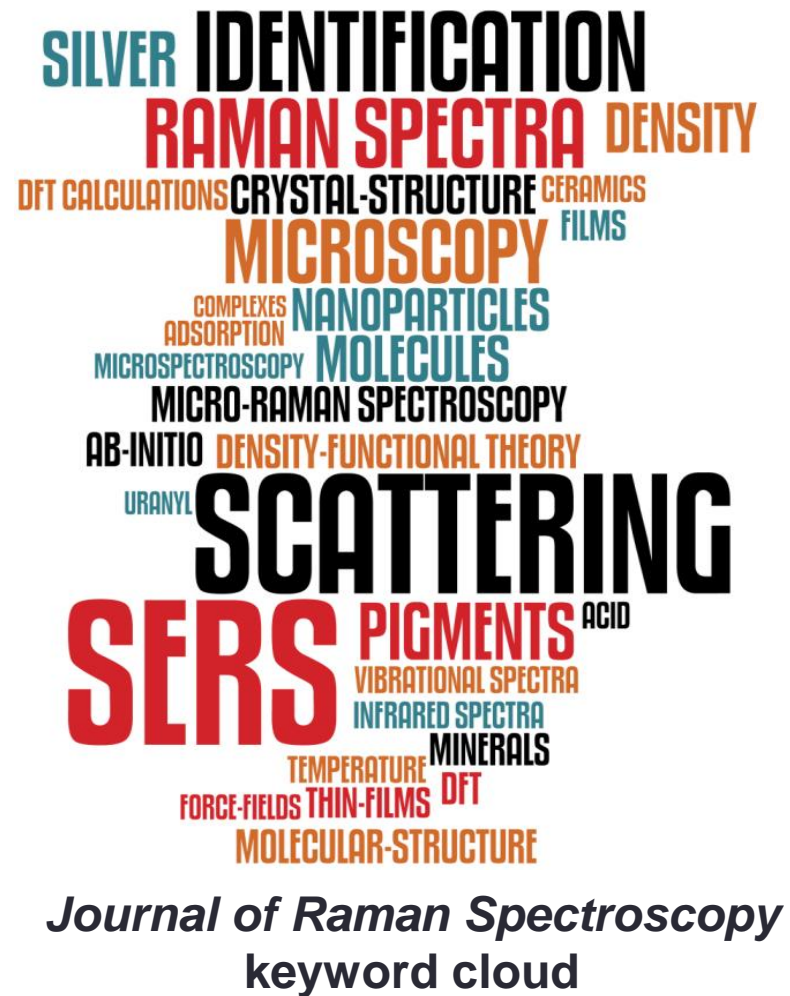
Hexachlorobenzene



Applications

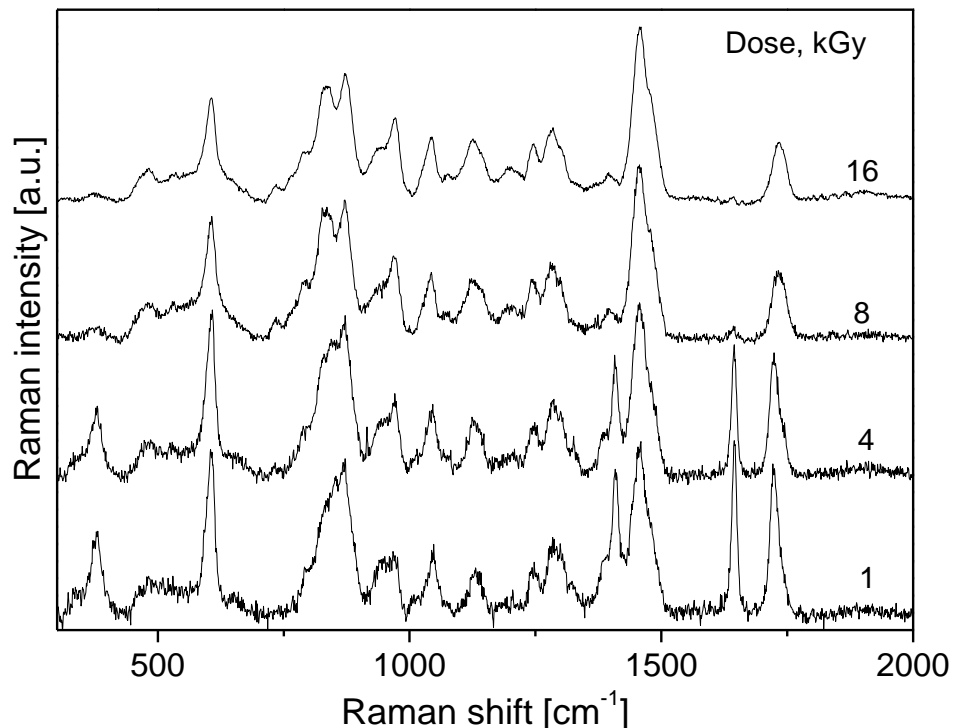
Information related to the vibrations of the medium

- Composition
- Conformation of the molecule
- Crystal lattice
- Crystal orientation
- Presence of isotopes
- Trace elements and defects
- Temperature
- Internal stress

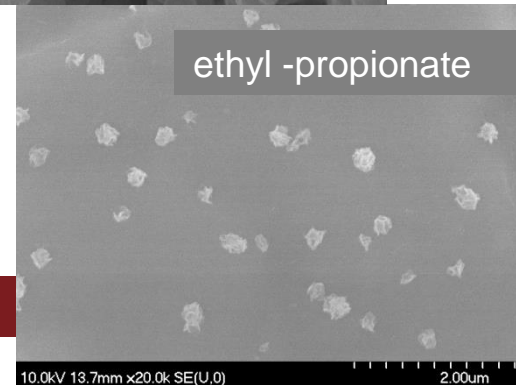
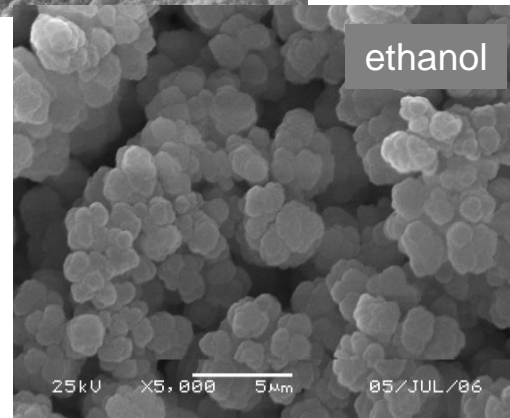
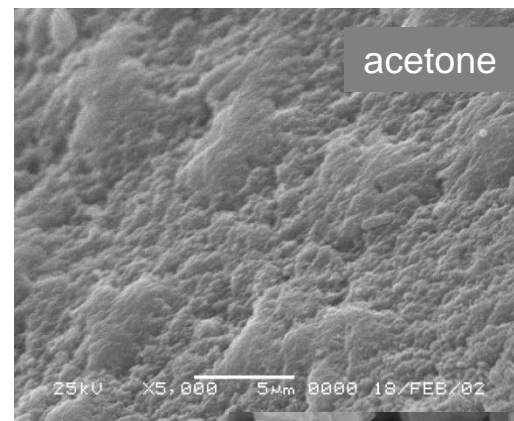


Study of polymerization kinetics

Gamma radiation induced precipitation polymerization of diethylene glycol dimethacrylate (DEGDMA) in different solvents.

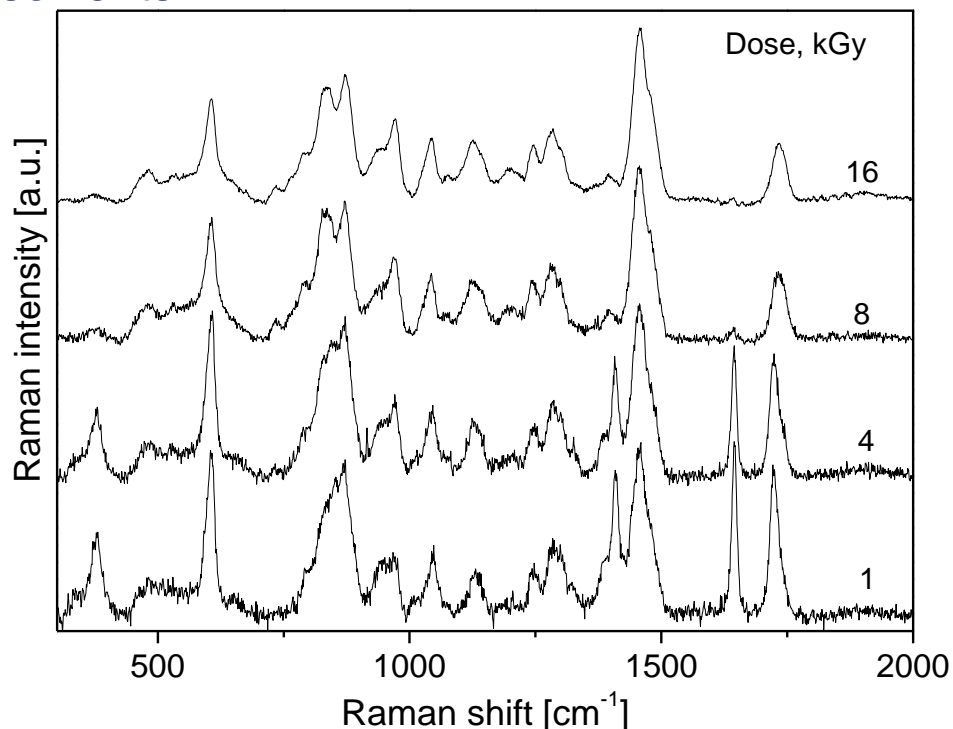


Raman spectrum of the monomer mixture polymerized with different doses.

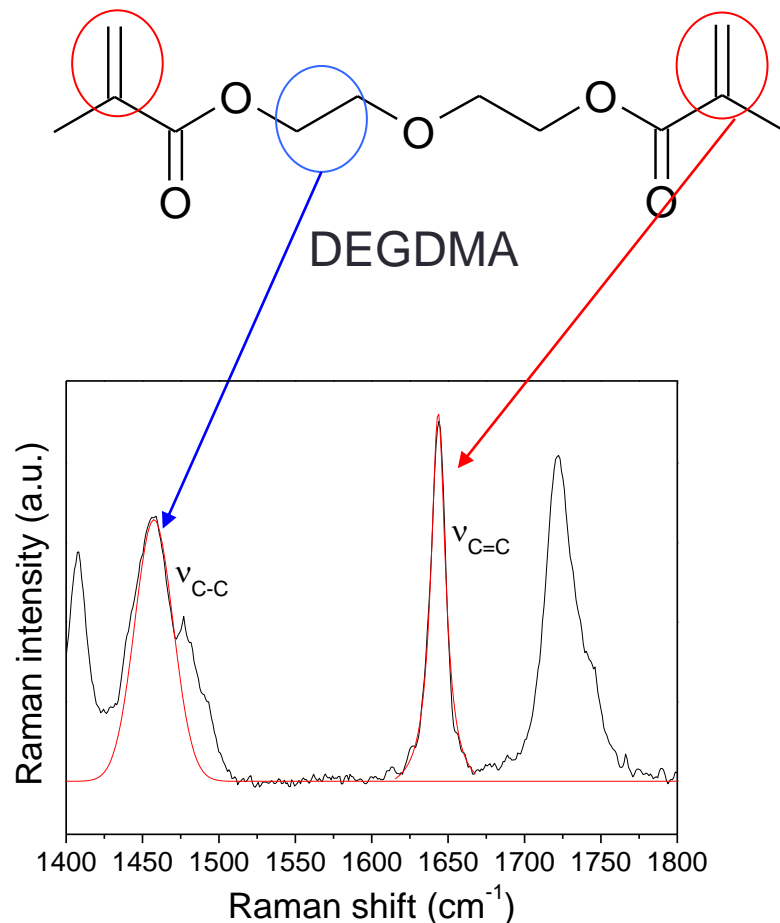


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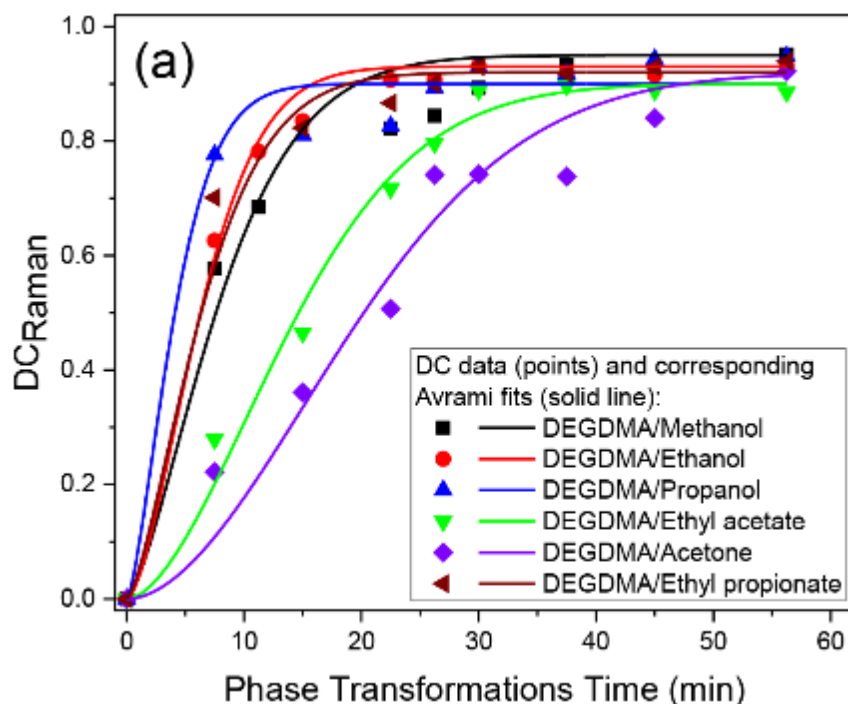
C=C vibration – 1640 cm^{-1}

C–C vibration – 1458 cm^{-1}

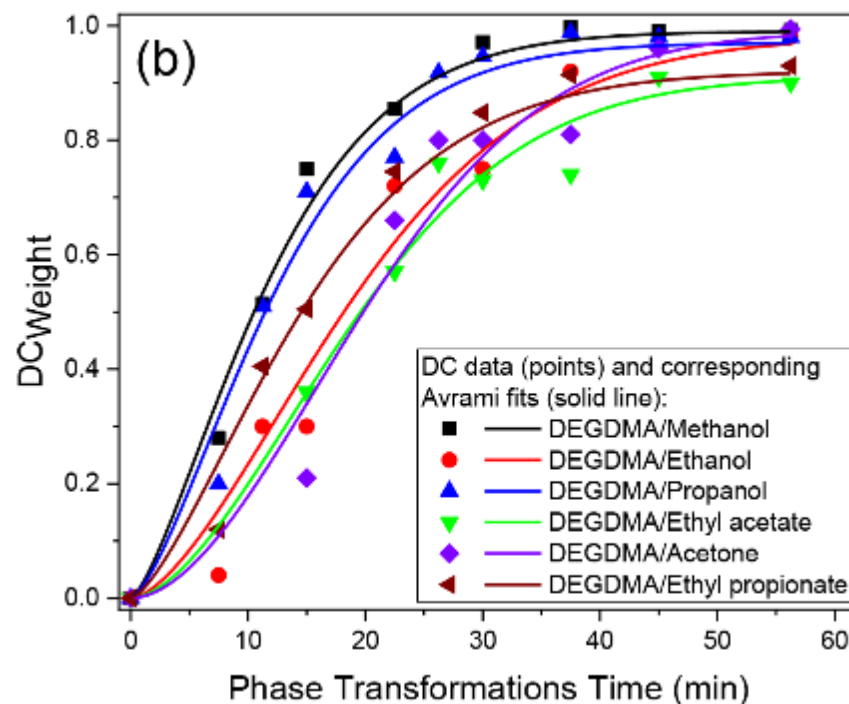
Study of polymerization kinetics

Conversion determined from Raman data and mass difference measurements.

Raman spectroscopy



Mass difference



- The C=C peak intensity >0 even for the maximum degree of conversion indicating that intact C=C bonds remain in the polymer matrix belonging to trapped or „dangling” monomers.

Polymer samples for irradiation experiments

Polymerization mixture

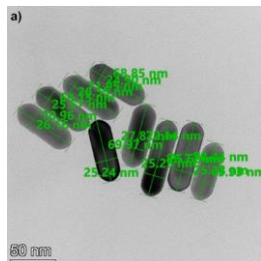
Monomers

- Urethane dimethacrylate (UDMA)
- Triethylene glycol dimethacrylate (TEGDMA)

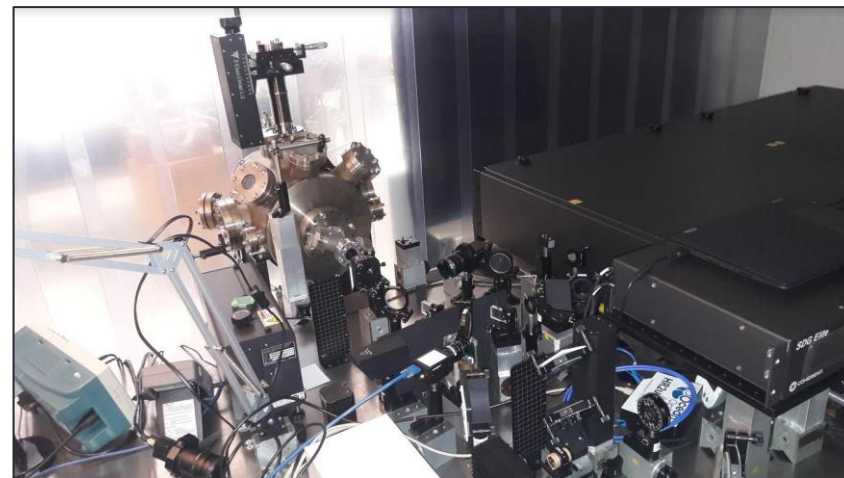
Photoinitiators

- Camphorquinone
- Ethyl 4-dimethylaminobenzoate

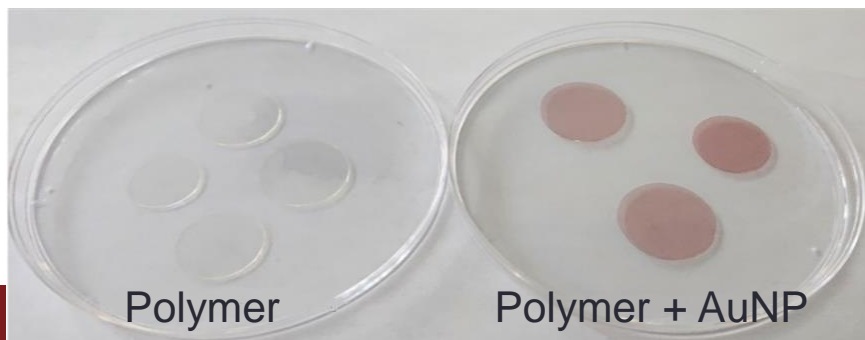
+ GOLD NANORODS (AuNP)



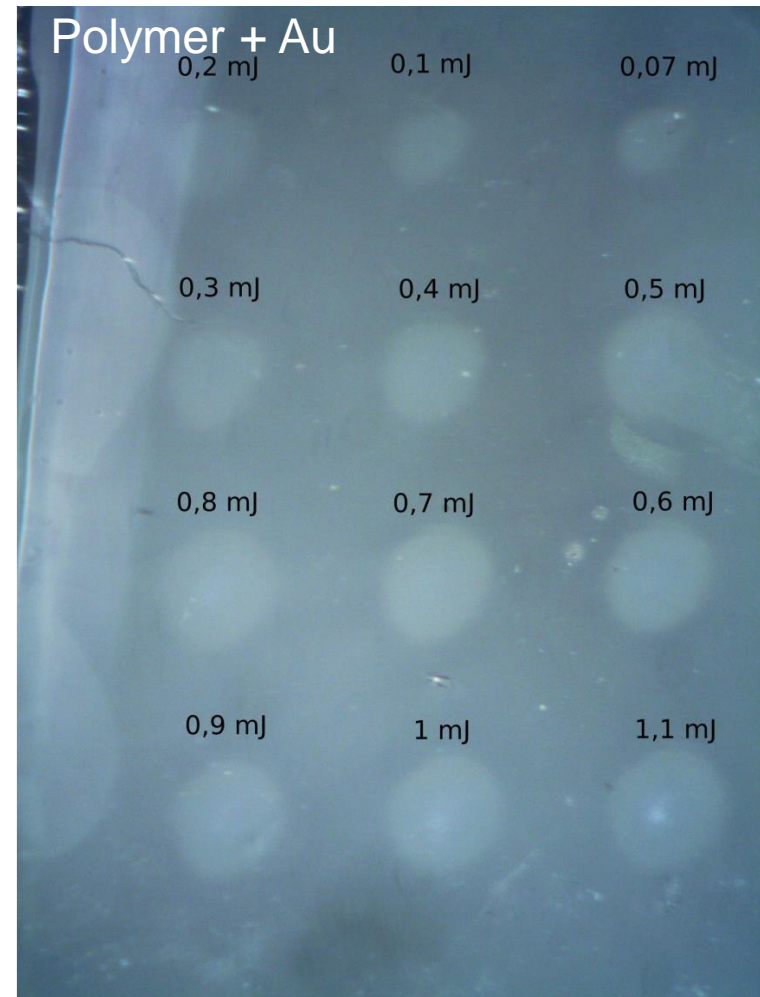
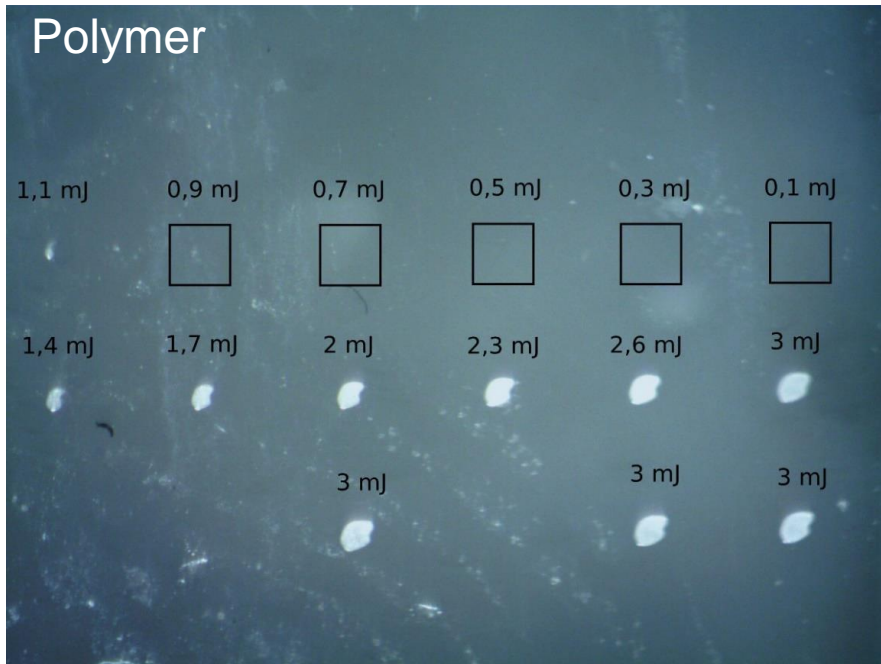
Irradiation



- Ti:Sa based chirped-pulse two-stage amplifier-laser system (Coherent Hydra)
- Smallest impulse length: 40 fs
- Central wavelength: 795 nm
- Repetition rate: 10 Hz
- Max. pulse energy: 30mJ
- Pressure in the vacuum chamber: $\sim 10^{-5}$ mbar



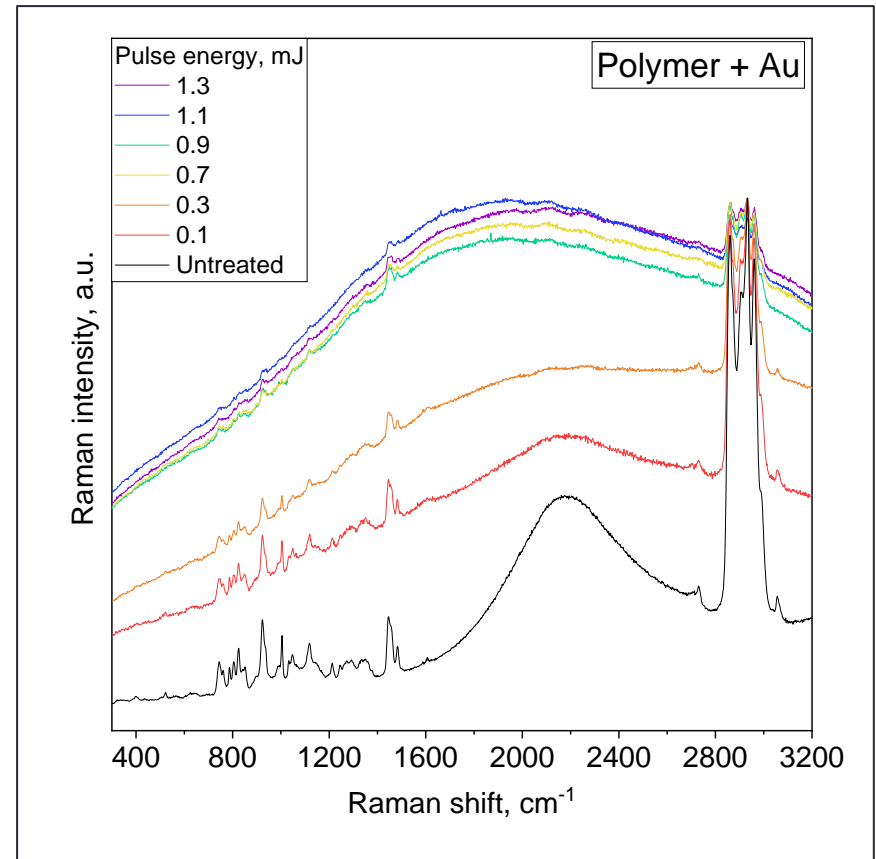
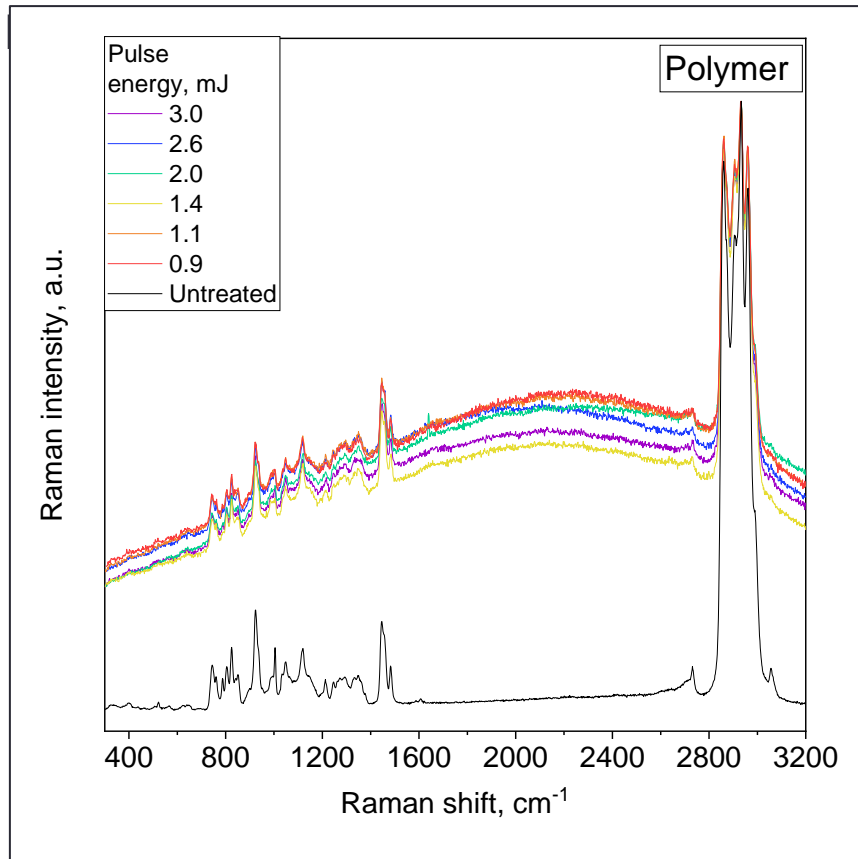
Low-energy laser pulses



- No crater formation
- The irradiation spot is visible from 1.1 mJ
- The spot size depends on the laser pulse energy

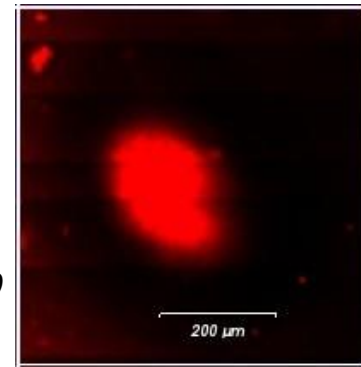
- The doping with gold nanorods affects the structural transformation of the polymer upon laser treatment

Low-energy laser pulses

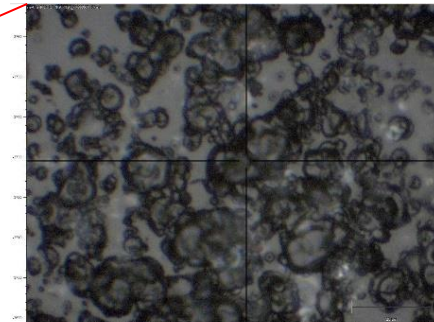
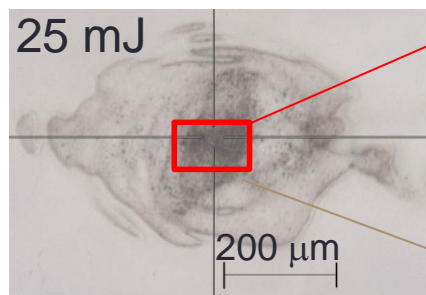
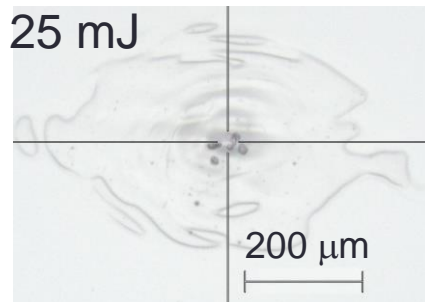
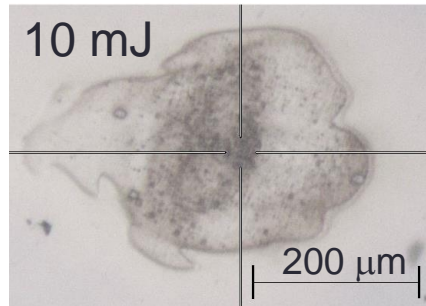
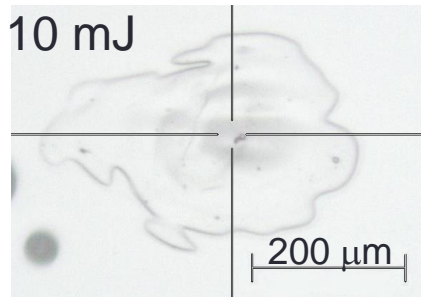
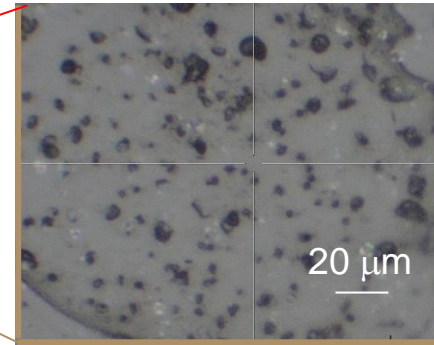
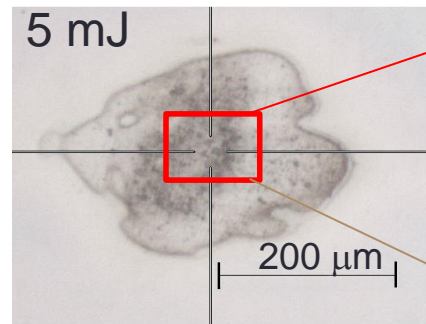
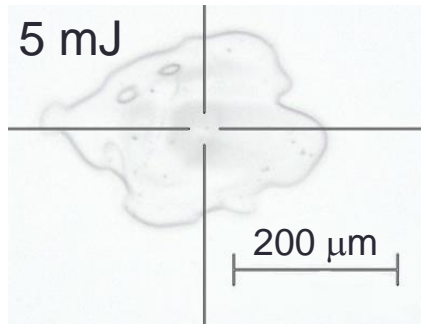


- The increase of the photoluminescence background is an indication of changes in the bonding configuration
- The Raman spectra of the non-doped polymer are very similar
- Gradual structural transformations occur in the polymer doped with gold nanorods

*Intensity map
at 1440 cm^{-1}*

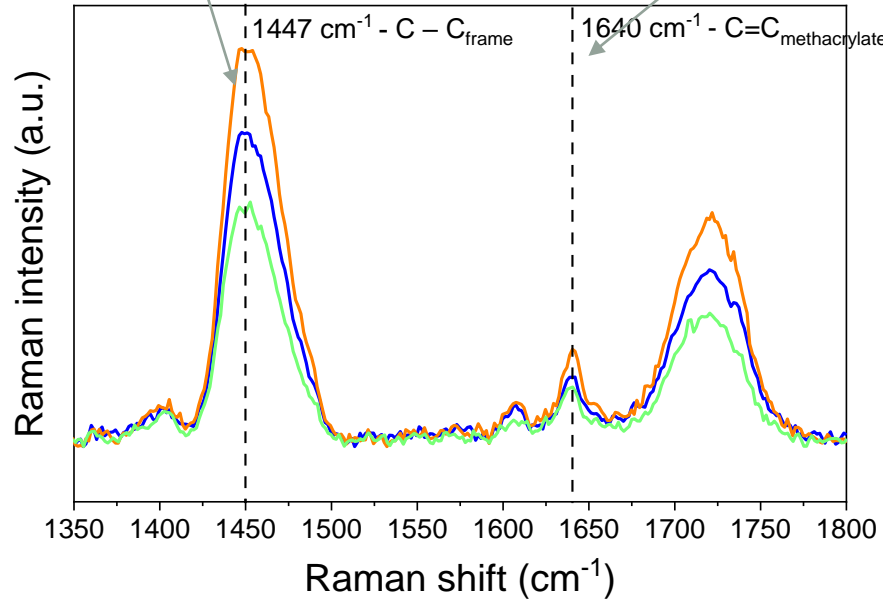
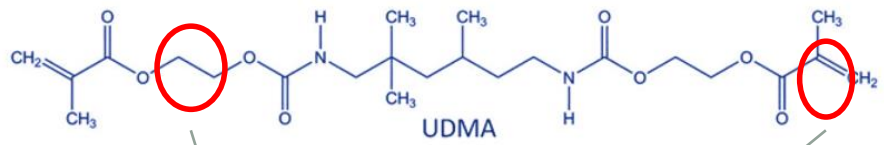


Higher pulse energies - Crater formation

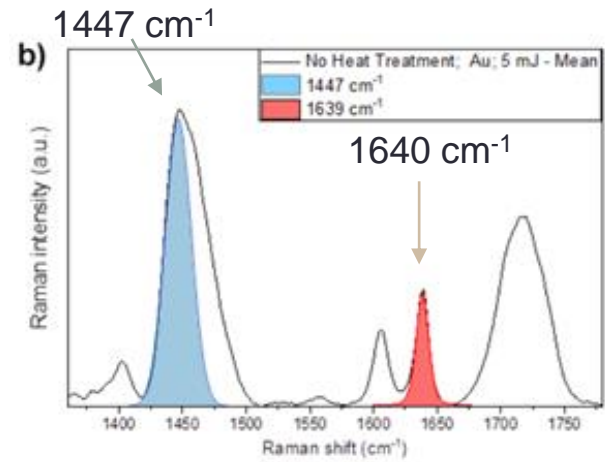
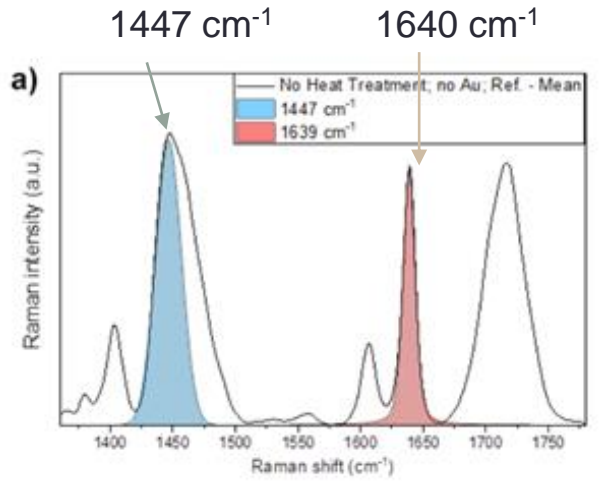
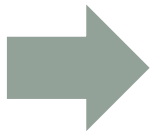


Raman characterization of the polymer

Structure of the urethane dimethacrylate (UDMA) monomer



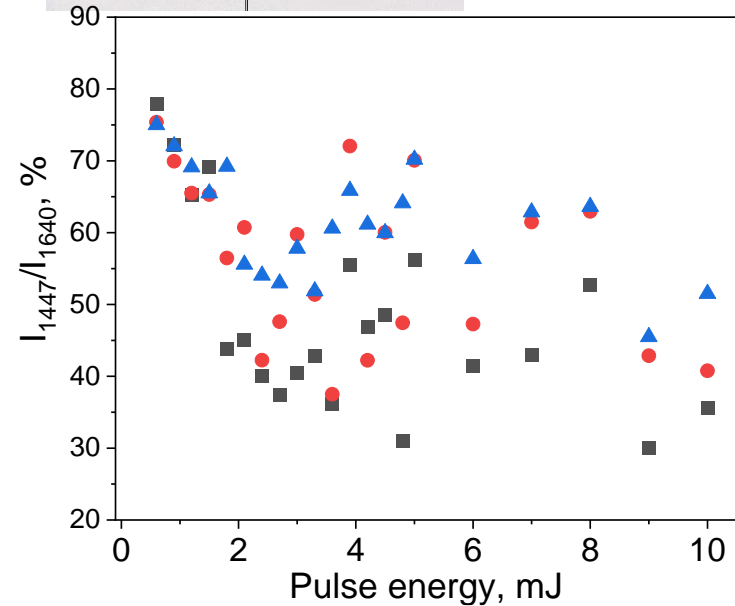
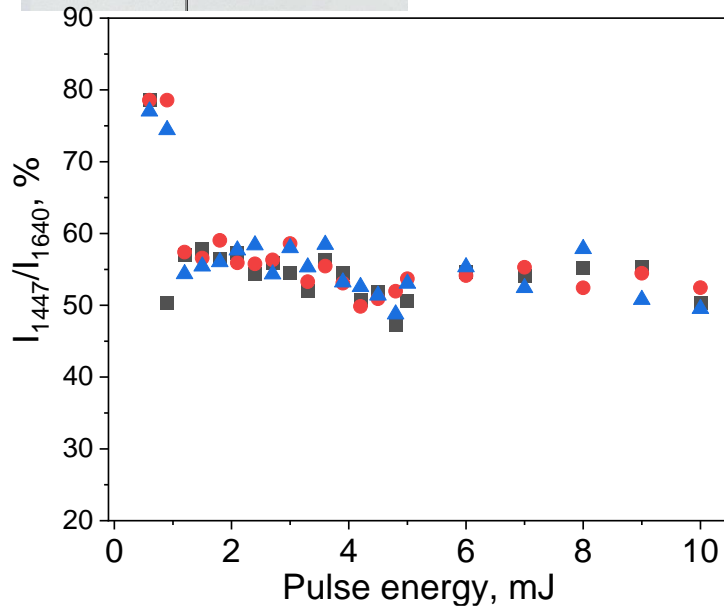
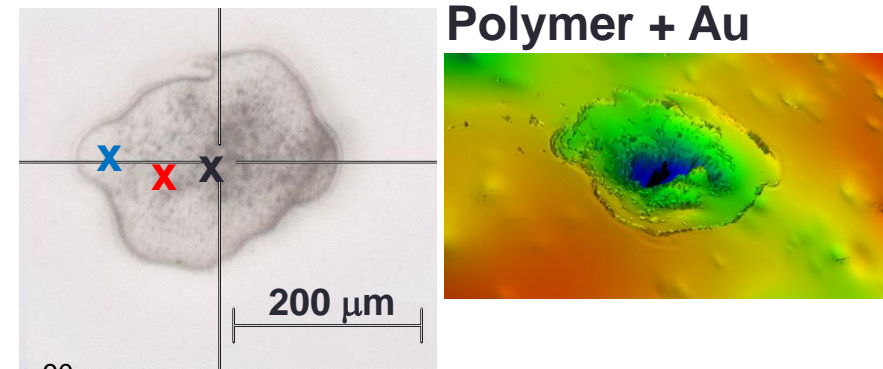
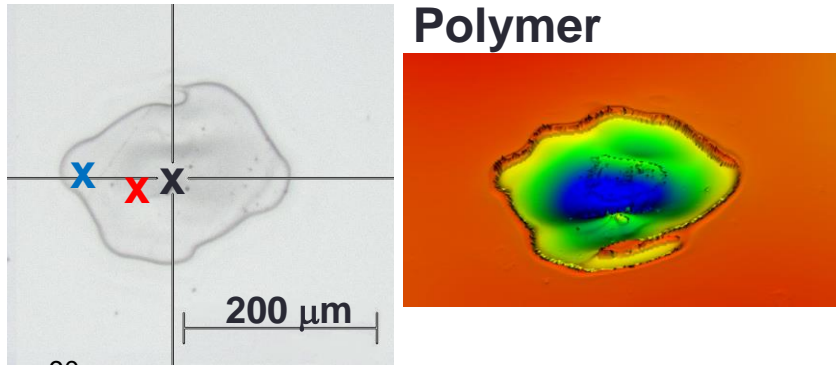
Typical Raman spectra of the UDMA polymer



During the polymerization, the ratio of the two peaks characterizes the degree of conversion.

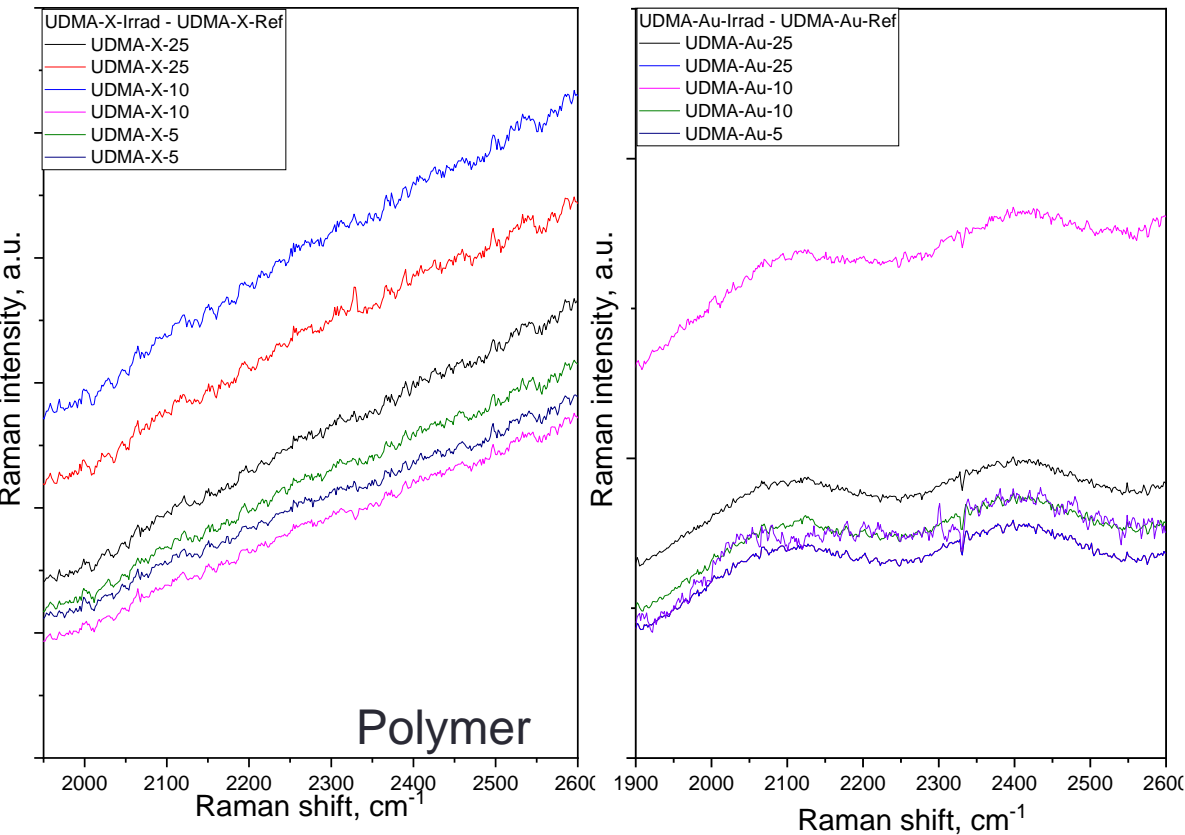
For the craters it shows the level of structural transformation – breaking of C-C bonds.

Higher pulse energies - Crater formation



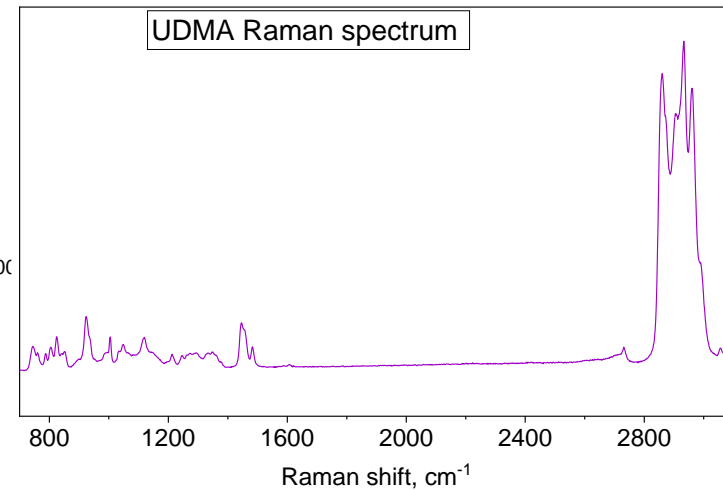
- The surface of the craters in non-doped polymer changes similarly, and that only slightly depends on the pulse energy.
- With gold nanorods the matrix is altered in a stochastic way, affected by the distribution of the nanorods. The change in the center is more remarkable.

Higher pulse energies - Crater formation



Possible origin:

- Photoluminescence
- Contamination
- Structural transformation

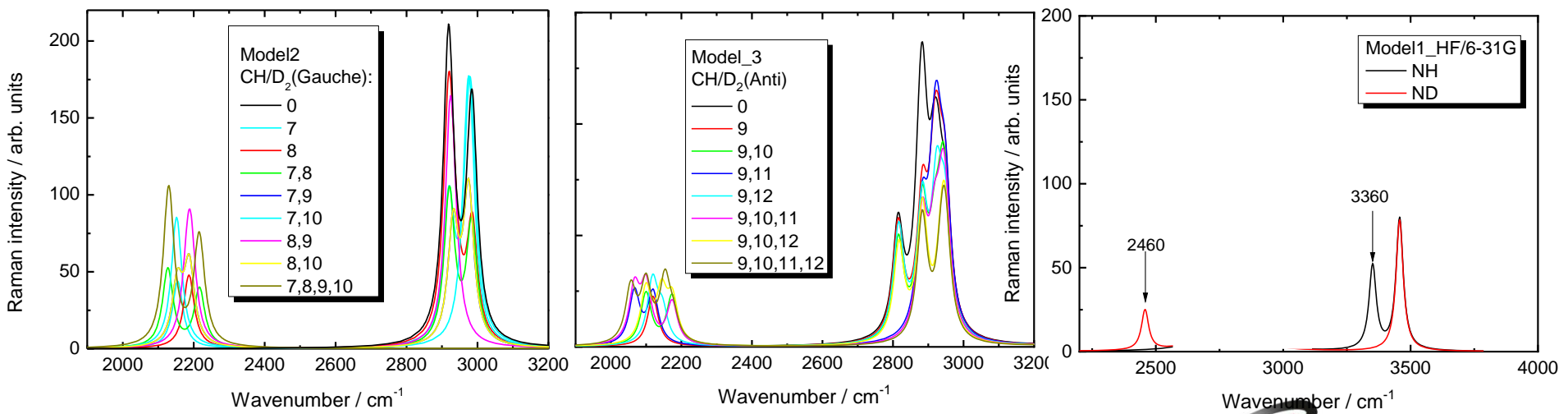
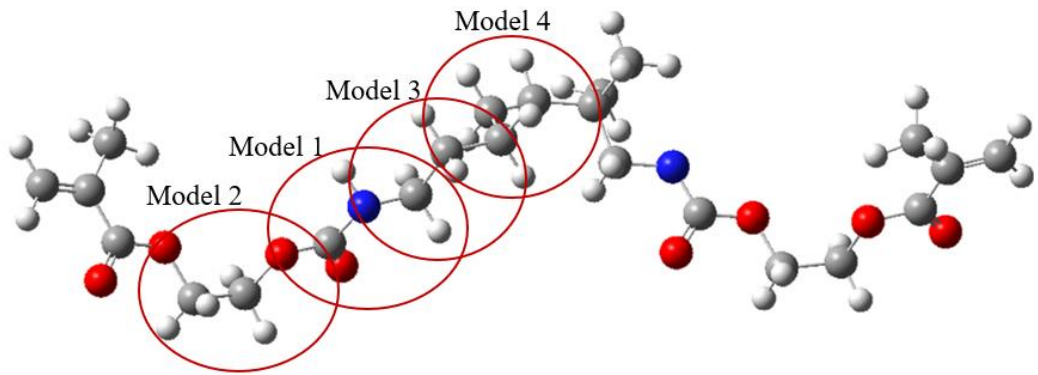


The Raman intensity of the polymer doped with gold increase remarkably in the 2000-2500 cm^{-1} region.

Origin of the new Raman peaks

Possible assignment – C-D vibrations

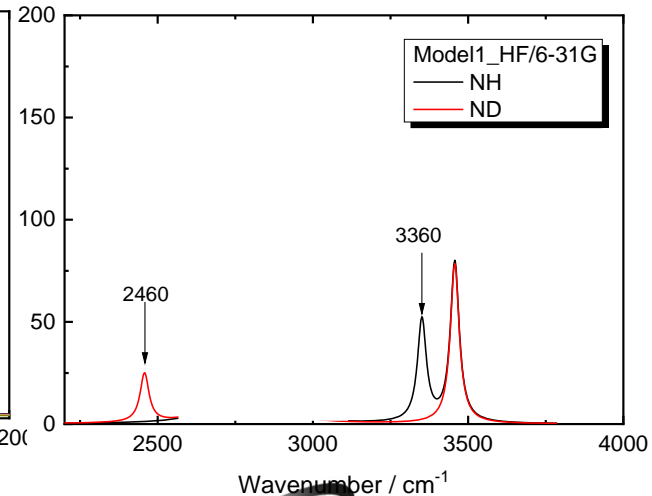
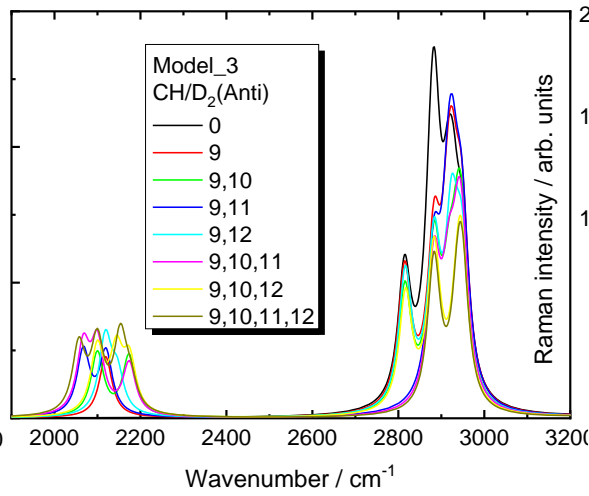
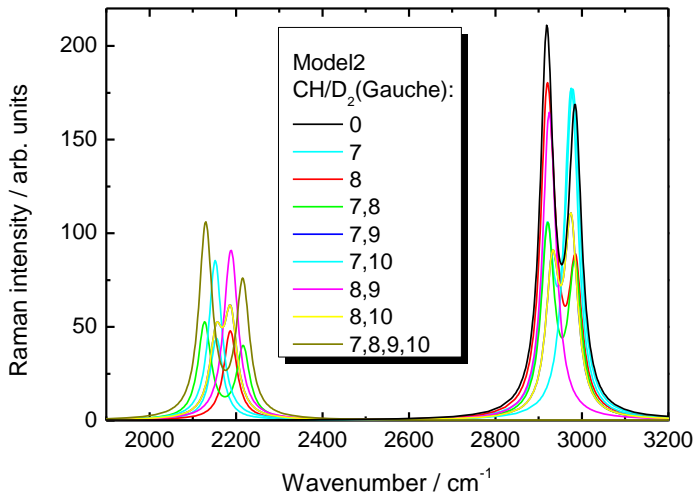
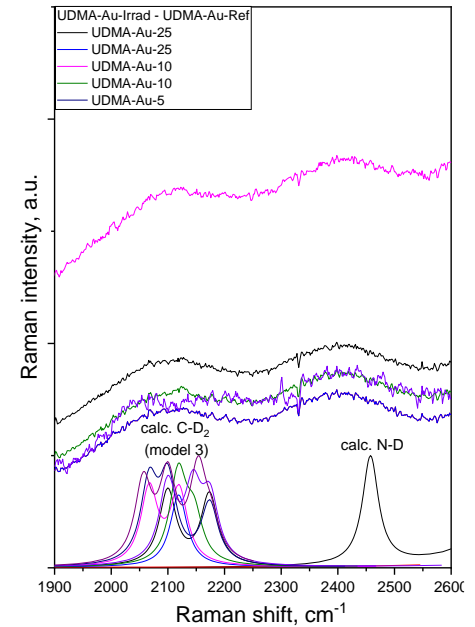
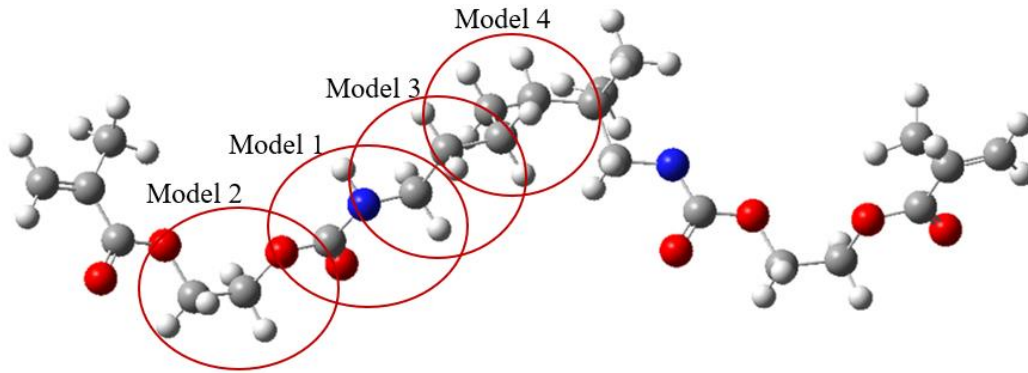
DFT calculations of Raman active vibrations of deuterized UDMA



Origin of the new Raman peaks

Possible assignment – C-D vibrations

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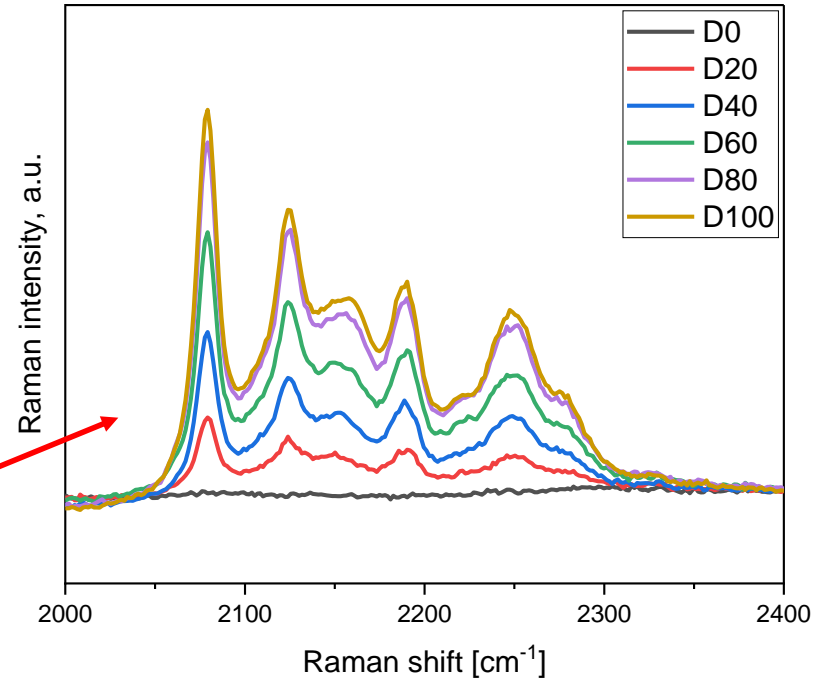


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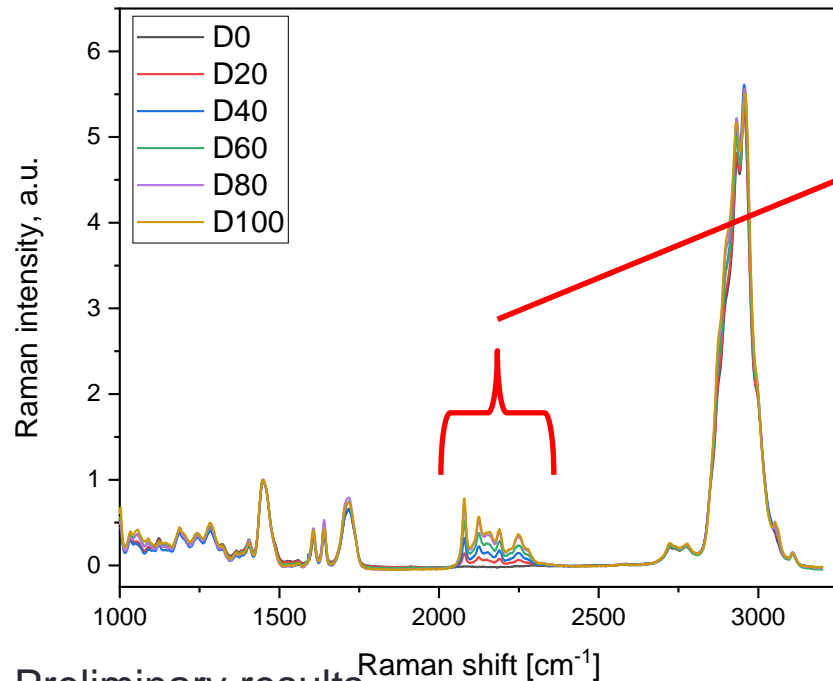
Possible assignment – C-D vibrations

Raman experiments with deuterized methyl methacrylate polymer of different deuterium content

Sample ID	Composition					
	MMA-D [g]	MMA [g]	UDMA [g]	Au [ml]	CQ [mg]	EDAB [mg]
D0	0	0,1	0,3	34,8	0,8	1,6
D20	0,02	0,08	0,3	34,8	0,8	1,6
D40	0,04	0,06	0,3	34,8	0,8	1,6
D60	0,06	0,04	0,3	34,8	0,8	1,6
D80	0,08	0,02	0,3	34,8	0,8	1,6
D100	0,1	0	0,3	34,8	0,8	1,6



The C-D vibrations give contribution in the 2050-2300 cm⁻¹ region of the Raman spectrum.



Preliminary results

Summary

- Raman spectroscopy is an optical spectroscopic technique allowing to study characteristic vibrations of the sample. It can be used to characterize the polymerization kinetics, the degree of conversion and structural transformations in polymers.
- The Raman spectroscopic study of polymer targets doped with gold nanorods and irradiated with ultrashort laser pulses showed that the presence of plasmonic gold nanoparticles has a remarkable effect on the structural transformations occurring due to light-matter interactions
- New Raman peaks were observed in the 2000-2500 cm^{-1} region of the Raman spectrum upon irradiation with high-energy laser pulses. The presumed origin of these features is the formation of C-D and N-D bonds in the structure.

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István Rigó



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ÉS INNOVÁCIÓS HIVATAL

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